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Explain giving reasons, which of the following sets of quantum numbers are not possible?
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~~Quantum numbers - allowable values identify which sets of quantum numbers are valid for an electron. each set is ordered (n, m, ms). Quantum numbers How many electrons in an atom could have these sets of quantum numbers Atomic Orbitals and Quantum Numbers Webinar: History of Atom Chemquest 12 Quantum Numbers Answers~~

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l – Secondary Quantum Number/Orbital Shape Quantum number: represents the shape of the orbital- s, p, f, d. l is a range of n-1. ml – Magnetic quantum number: represents the number of orbits possible. M l is a range of l. ms – Spin Quantum number: represents the electron and its spin. Two possibilities +1/2,

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When l = 0 (s cloud), 1 (p cloud), 2 (d cloud), 3 (f cloud) Third – Magnetic Quantum number (ml) =

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location or spatial orientation $m_l = -l$ to l Fourth – Spin Quantum number (m_s) = electron spin $m_s = +1/2$ or $-1/2$. 2. Name the orbitals described by the following quantum numbers a. $n = 3, l = 0$ 3s. c. $n = 3, l = 2$. 3d

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